Introduction to General Chemistry CHM 1025C   **Instructor: Mr. George W.J. Kenney, Jr**
CT #1, Ch 1, 2, 3   **22-Sept-2008 West**

NO CREDIT IF YOU: Fail to put in the Units & Properly Round, Fail to show ALL math work,

PRINT YOUR NAME on the line: _____
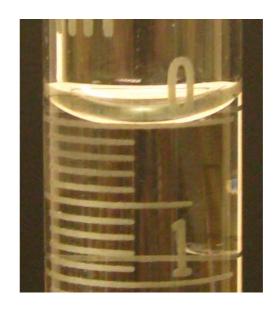
Your start time on this test   _____

Your finish time on this test: _____

Time it took you to do this test: _____

1. (30 pts, 3 pts ea) FILL IN THE BLANKS

1.1 The Scientific Approach utilizes   _____ ,

1.2   _____

1.3   _____

1.4 An Ethyl Alcohol Water mixture can be separated into it's separate components by:

_____

1.5 Scientific Notation deals with: _____

1.6   _____

1.7 Which temperature scale is used in most chem 1025 laboratory experiments?

_____

1.8 Give an example of a Physical Property: _____

1.9 Give an example of a Chemical Property: _____

1.10 Give an example of a Heterogeneous Mixture: : _____

2. (32 pts, 10 pts ea) Perform the following math conversions SHOW ALL MATH

2.1 Convert -17. ^oF to K?

2.2 Convert 234.6 mg to Tons

2.3 Convert 55 miles to km

2.4 **(8 pts, 4 pts ea)**   Read the following burettes:

_____   _____

3. (30 pts, 5 pts ea) Perform the following math operations:

3.1 123456789.23 cm to 2 significant digits: _____

3.2 2.000 g * 1.2 in =   _____

3.3 Perform the following addition operations directly as shown to calculate the molecular weights:

    $C_2H_6O_2$ C 2 * 12.01 =   24.02   g mole
                   H 6 * 1.008 =   6.048 g/mole
                   H 2 * 16.01 =   <u>32.02   g/mole</u>

3.4   HCl   H 1 * 1.008 =   1.008 g/mole
                   Cl 1 * 35.45   <u>35.45</u>

3.5 32.0 ^{o}F is equal to _____ ^{o}C?

3.6 How many significant digits are in 00010203040.506070 grams? _____

How do you rate this test from 1 to 10   _____
1 = Very East, can do it with my eyes closed, 10= Very Very Difficult, could not do any of the problems